

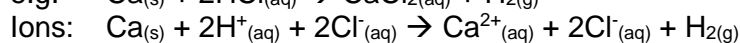
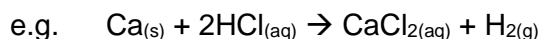
Predicting compounds

1. Write a **balanced** equation for the following acid-metal reactions.

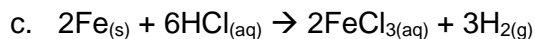
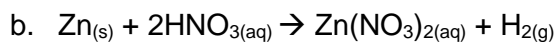
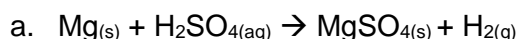
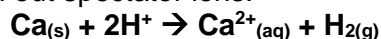
Include state symbols.

- Solid calcium sulphate is produced from a reaction.
- Aqueous zinc chloride is produced.
- A clear, colourless liquid containing magnesium sulphate is produced.
- An opaque white liquid containing lead chloride is produced.

2. Convert the following equations into ionic equations:



Cancel out spectator ions:



- d. A generic group 2 metal reacting with a generic acid

3. For each case above, show what is being oxidised and what is being reduced. Write a half equation for each transformation. (See Y9 work)

